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**7.B Mole and Mass Avogadro's Number, The Mole, Grams, Atoms, Molar Mass Calculations - Introduction** Converting Between Grams and Moles Mole Ratio Practice Problems **Concept of Mole - Part 1 | Atoms and Molecules | Don't Memorise** Empirical Formula \u0026amp; Molecular Formula Determination From Percent Composition **Molality Practice Problems - Molarity, Mass Percent, and Density of Solution Examples** ~~Introduction to Moles~~ Stoichiometry Basic Introduction, Mole to Mole, Grams to Grams, Mole Ratio Practice Problems *Mole Conversions Made Easy: How to*

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*Convert Between Grams and Moles How To Convert Grams To Moles - VERY EASY! Counting Atoms: Intro to Moles Part 2 ~~Converting Grams to Moles Using Molar Mass | How to Pass Chemistry The Mole: Avogadro's Number and Stoichiometry~~ Introduction to Combustion Analysis, Empirical Formula & Molecular Formula Problems Dalton's Law of Partial Pressure Problems & Examples - Chemistry Stoichiometry Mole to Mole Conversions - Molar Ratio Practice Problems Step by Step Stoichiometry Practice Problems | How to Pass Chemistry ~~Balancing Chemical Equations Practice Problems~~ Concept of Mole + Avogadro's Number + Atoms and Molecules + ~~Don't Memorise~~ *Chemistry B Moles Packet Answer**

CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 9 CHEMISTRY WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass,  $6.02 \times 10^{23}$  molecules and, for a gas, 22.4 liters at STP.

CHEMISTRY B- MOLES PACKET NAME: HR: \_\_\_\_\_ PAGE 1

CHEMISTRY B- MOLES PACKET NAME: \_\_\_\_\_ HR: \_\_\_\_\_ PAGE 4 CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.

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A mole of carbon contains  $6.02 \times 10^{23}$  atoms of carbon. A mole of helium contains  $6.02 \times 10^{23}$  atoms of helium. A mole of sodium contains  $6.02 \times 10^{23}$  atoms of sodium. A mole of gold contains  $6.02 \times 10^{23}$  atoms of gold. However, we have learned about seven elements that exist as diatomic molecules— $H_2$ ,  $N_2$ ,  $O_2$ ,  $F_2$ ,  $Cl_2$ ,  $Br_2$ , and  $I_2$

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CHEMISTRY WORKSHEET # 2: THE MOLE AS A UNIT OF MASS  
Now that you know how to find the mass of one mole of a substance (molar mass) you can easily find the mass of several moles or the mass of a fraction of a mole using the factor-label technique.  
 $1 \text{ mol} = \text{a molar mass of an atom/molecule (g/mol)}$   
EXAMPLE: What is the mass of 5.00 moles of water  $\text{H}_2\text{O}$ ?  
STEP 1:  $2 \text{ H} = 2 \times (1.01) = 2.02$   
STEP 2:  $? \text{ grams H}_2\text{O} = 5.00 \text{ moles H}_2\text{O} \times 18.02 \text{ g H}_2\text{O} = 90.10 \dots$

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CHEMISTRY WORKSHEET # 4 MOLAR MASSES AND AVOGADRO'S NUMBER  
Now that you know two definitions of a mole (a molar mass and an Avogadro's number of particles) you can combine these two definitions into one problem.  
EXAMPLE: How many molecules are there in 90.1 grams of water?  
 $2 \text{ H} = 2 \times (1.01) =$

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$2.020 = 1 \times (16.00) = 16.00$  H<sub>2</sub>O = 18.02 g?  
molecules H<sub>2</sub>O = 90.1 g H<sub>2</sub>O  $\times$  1 mole H<sub>2</sub>O  
 $\times 6.02 \times 10^{23}$  molecules H<sub>2</sub>O =  $3.01 \times 10^{24}$   
molecules H<sub>2</sub> ...

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WORKSHEET # 6 MIXED MOLE PROBLEMS (GRAMS, MOLECULES, AND LITERS) You now know three things a mole can be: a molar mass,  $6.02 \times 10^{23}$  molecules and, for a gas, 22.4 liters at STP. We can use this

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Question: PAGES CHEMISTRY B-MOLES PACKET

NAME: HR: CHEMISTRY WORKSHEET #5 MOLE

PROBLEMS-MOLAR VOLUME OF A GAS We Have

Learned That A Mole Is A Mass Of Material And

Number Of Particles. A Mole Can Also Be A

Measure Of Volume When We Are Talking About

Gases. You May Remember From Previous Science

Classes That All Gases Behave Basically The

Same As Far As The Physical ...

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