

Bond Energy Pogil Answer

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Use the bond energies in Model 2 to calculate the energy that is released to form the two moles of molecules in Step 2 of the reaction in Model 3. Include the proper sign and units on your answer. The heat of reaction or change in enthalpy (ΔH) of a reaction is the net energy change in the reaction. Consider where ΔH is located in Model 3.

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9. Is the bond energy for a double bond simply two times that of a single bond for the same atoms? Provide two specific examples to support your answer. 10. Explain how to use the information in Model 2 to determine the change in energy when a mole of C—H bonds is formed rather than broken. 11. Because the environment surrounding a bond in a ...

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The bond enew for a particular type of bond can vary from one molecule to another because the atomic environment of a bond can influence the amount of energy needed to break the bond. For example the carbon—carbon bond in the two molecules shown below may not have the same bond enerW because the surrounding atoms are different.

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436. Cl - Cl. 243. H - Cl. 432. Energy in = 436 + 243 = 679 kJ/mol. Energy out = 2 x 432 = 864 kJ/mol. Energy change = in - out = 679 - 864 = -185 kJ/mol. The energy change is negative ...

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